

Corrosion products and mechanism on NiTi shape memory alloy in physiological environment

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Despite many investigations on the corrosion behavior of NiTi shape memory alloys (SMAs) in various simulated physiological solutions by electrochemical measurements, few have reported detailed information on the corrosion products. In the present study, the structure and composition of the corrosion products on NiTi SMAs immersed in a 0.9% NaCl physiological solution are systematically investigated by scanning electron microscopy (SEM), x-ray energy dispersion spectroscopy (EDS), and x-ray photoelectron spectroscopy (XPS). It is found that attack by Cl^- results in nickel being released into the solution and decrease in the local nickel concentration at the pitting sites. The remaining Ti reacts with dissolved oxygen from the solution to form titanium oxides. After long-term immersion, the corrosion product layer expands over the entire surface and XPS reveals that the layer is composed of TiO_2 , Ti_2O_3 , and TiO with relatively depleted Ni. The growth rate of the corrosion product layer decreases with immersion time, and the corrosion product layer is believed to impede further corrosion and improve the biocompatibility of NiTi alloy in a physiological environment. It is found that the release rate of nickel is related to the surface structure of the corrosion product layer and immersion time. A corrosion mechanism is proposed to explain the observed results.

I. INTRODUCTION

NiTi SMAs as biomaterials are used in medical implants and devices such as orthodontic wires, self-

expanding cardiovascular and urological stents, spine correction rods, bone fraction fixation plate and staples, and so on.¹⁻⁵ The reasons for adopting NiTi SMAs in biomedical implants are their unique shape memory effects and superelasticity properties, low Young's modulus compared with stainless steels and titanium alloys, reliable and stable mechanical properties, as well as good biocompatibility.⁶⁻⁸ However, as are other metallic implants, when

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NiTi SMAs are subjected to the physiological environment, the corrosion behavior affects not only their biocompatibility but also long-term implantation safety in the human body. It has been reported that corrosion causes outdiffusion of harmful nickel ions that may induce hypersensitive reactions and tissue necrosis.^{9,10} Because the human body constitutes an aggressive corrosion environment, the presence of ionic salts in body fluids can corrode NiTi alloys and further accelerate Ni outdiffusion from the substrate. Up to now, although many studies have been carried out to investigate the corrosion behavior of biomedical NiTi alloys under different physiological environments *in vitro*, most of them have focused on Ni release from the bulk materials to the physiological solution and electrochemical properties such as the open circuit potential, dynamic polarization as well as electrochemical impedance spectroscopy (EIS), which have been considered to be closely linked to biocompatibility.^{11–19} After the NiTi alloys are implanted into human body, they will be attacked by aggressive ions in the body fluids such as chlorides, hydrocarbonates, and sulfates that further corrode the NiTi implants and form corrosion products on the surface. The nature of this corrosion product layer not only influences the subsequent corrosion process of the substrate,⁸ but may also affect the biological behavior of bone tissues. However, the structure of the corrosion product layer, exact constituents and their distribution, and formation mechanisms are still not clear and well understood. The main reason is that the corrosion rate of NiTi SMAs in a physiological solution is low compared with other biomedical metals such as magnesium alloys.²⁰ Therefore, it is crucial to identify the constituents and mechanism of the corrosion product layer on NiTi SMAs to enable more trouble-free clinical use of the materials.

Because human body fluids contain a high concentration of Cl^- that can cause pitting corrosion on the surface of implants,²¹ the present work aims at investigating the corrosion products and mechanism on NiTi SMAs in an aggressive chloride ion environment. In the work described here, immersion tests are conducted in 0.9% NaCl physiological solution. The morphology and composition of the NiTi surface are investigated by scanning electron microscopy (SEM), x-ray energy dispersion spectroscopy (EDS), and x-ray photoelectron spectroscopy (XPS). The amounts of nickel ions released from the NiTi specimens were measured by inductively coupled plasma atomic emission spectrometer (ICPAES). The corresponding corrosion mechanism of the NiTi SMAs in 0.9% NaCl physiological solution is proposed.

II. EXPERIMENTAL DETAILS

The materials used were commercial superelastic NiTi bar (SE508, Nitinol Device & Components, Fremont, CA) consisting of 50.8 at.% Ni and 49.2 at.% Ti. The

raw materials were cut into 1 mm thick and 4.8 mm diameter disks. The specimens were polished by SiC sandpaper to grade 1200 followed by ultrasonic cleaning in acetone and deionized water for 10 min. The cleaned specimens were immersed in 80 mL of 0.9% NaCl physiological solution in a polypropylene (PP) bottle. The polypropylene bottles were closed tightly and incubated in a thermostatic chamber at 37 ± 0.5 °C for 2, 5, 8, 12, and 30 weeks, respectively. Each PP bottle contained one NiTi specimen, and there were three parallel PP bottles for each time point. After different immersion times, the solutions were analyzed by ICPAES to determine the amounts of Ni leached from the specimens.

The surface morphologies on the NiTi specimens were observed using SEM. The surface composition and depth profiles in the near-surface region of the NiTi samples were determined by XPS (Physical Electronics PHI 5802, Chanhassen, MN) with monochromatic $\text{Al K}\alpha$ x-ray radiation (1486.6 eV). High-resolution Ti 2p, O 1s, and Ni 2p spectra were acquired at 20 eV pass energy to determine the chemical states and concentrations. The C1s peak was used to identify any charging effects. The XPS depth profiles were obtained using a rastered 3 keV Ar^+ ion beam. The oxide thickness was estimated by taking the depth when the O signal dropped to 50% of the maximum value. The sputtering rate of the oxide film was estimated to be 7.8 nm/min based on an average sputtering rate calculated from a SiO_2 reference under similar conditions.²²

III. RESULTS AND DISCUSSION

Figure 1 shows the SEM images acquired from the NiTi specimens immersed for 2, 5, and 8 weeks. As shown in Fig. 1(a), some dark spots observed on the surface of the NiTi specimen after immersion for 2 weeks can be identified as initial pitting corrosion in the NaCl solution. Figure 1(b) depicts more serious pitting corrosion after immersion for 5 weeks. With attack by Cl^- from the NaCl solution, the dark spots transform into tiny pores on the surface. When the immersion time is increased to 8 weeks, some of the pores propagate into the large corrosion spots as indicated by arrows in Fig. 1(c). The high-magnification image shows that these corrosion spots feature a hole in the center surrounded by several gray spots indicated by white arrows in Fig. 1(d). These newly formed gray spots indicate that after immersion for 8 weeks, new pitting sites still occur on the surface of NiTi specimens. According to the color variation observed from the corrosion spots, it can be concluded that the darker the corrosion spot, the more severe is corrosion. Pitting corrosion causes dissolution of the bulk materials into the corrosive environment, consequently forming pores on the NiTi SMAs.

To determine the elemental compositions in the corroded layer, four representative regions 1 to 4 in Fig. 1(d)

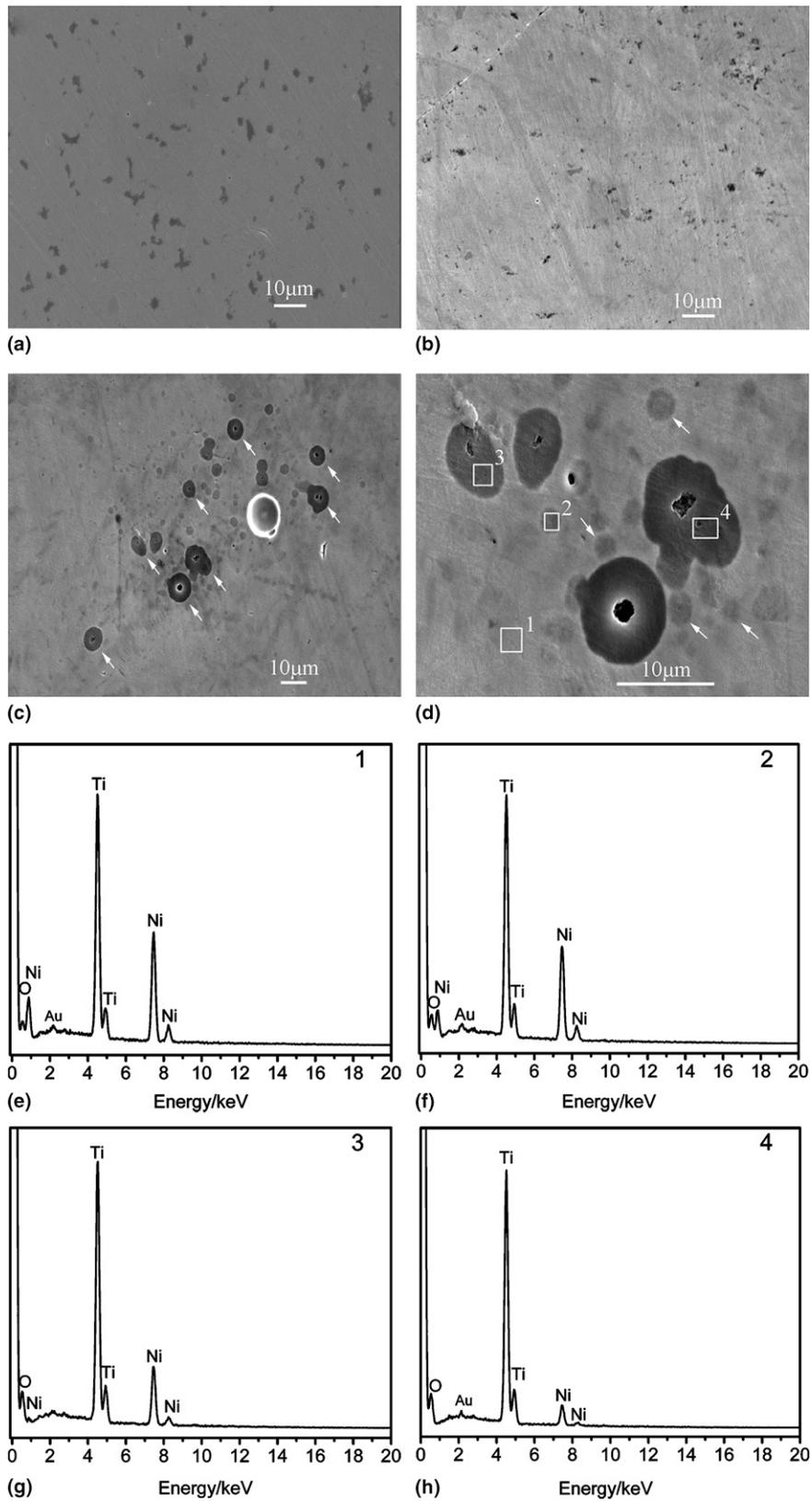
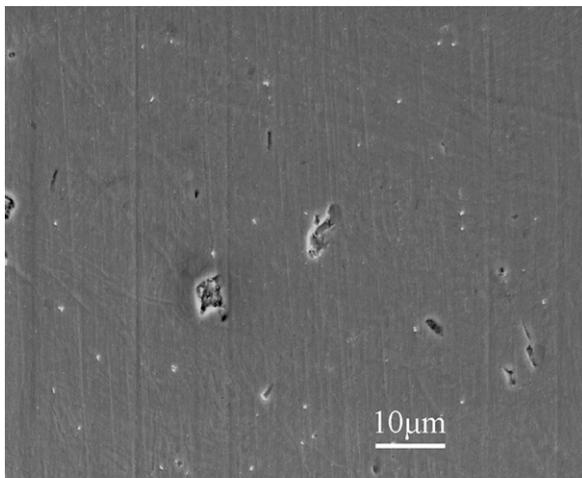


FIG. 1. Surface morphology of NiTi specimens after immersion in the 0.9% NaCl physiological solution after: (a) 2 weeks, (b) 5 weeks, and (c) 8 weeks; (d) high magnification of (c); (e)–(h) EDS spectra acquired from selected regions 1–4 in (d).

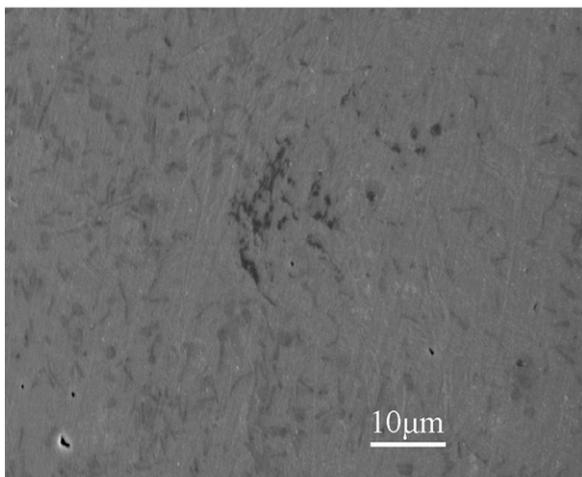
were analyzed by EDS after immersion for 8 weeks, and the results are displayed in Figs. 1(e) to 1(h). It can be found that Region 1 with slight corrosion has a relatively high concentration of Ni. In comparison, surface corrosion becomes severe in Regions 2 to 4 with the surface concentration of Ni decreasing significantly and those of O and Ti increasing gradually. The detailed elemental concentrations are listed in Table I. The low Ni concentration on the corroded surface suggests that pitting corrosion results in nickel release from the bulk materials to the NaCl solu-

TABLE I. Elemental composition of NiTi specimen immersed in the 0.9% NaCl physiological solution for 8 weeks.

Region	Element composition (at.%)		
	O	Ti	Ni
d1	62.33	15.69	21.98
d2	62.2	20.80	16.48
d3	69.19	21.09	9.61
d4	74.06	22.18	3.62



(a)

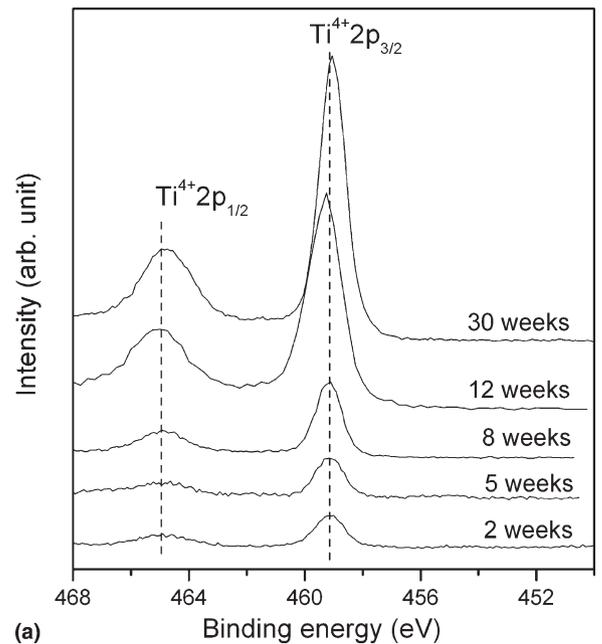


(b)

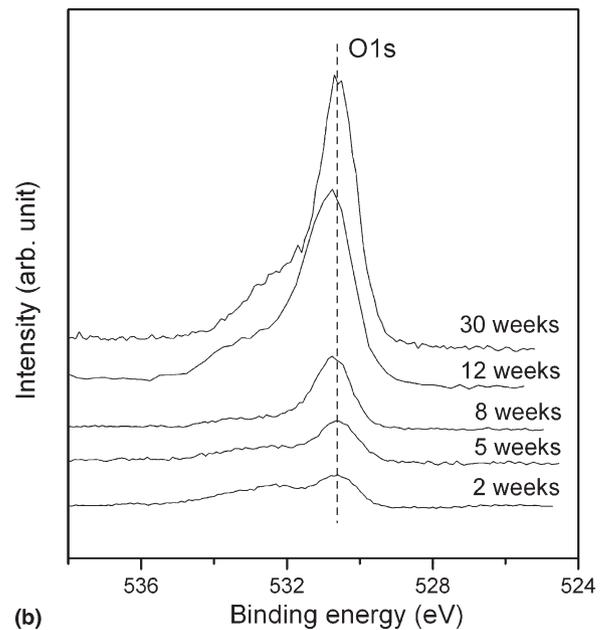
FIG. 2. Surface morphology of the NiTi specimens after immersion in 0.9% NaCl physiological solution for (a) 12 weeks and (b) 30 weeks.

tion. Meanwhile, the remaining Ti reacts with dissolved oxygen in solution to form titanium oxides in the corroded area. The atomic ratio of O to Ti is greater than 2 attributable to chemical absorption of $-\text{CO}$ on the surface. Our results demonstrate that nickel release and titanium oxide formation occur simultaneously during pitting corrosion on NiTi SMAs in the NaCl solution. As a result, the corrosion product consists of a layer of titanium oxides.

Figures 2(a) and 2(b) show the typical surface morphologies after immersion in the 0.9% NaCl solution for 12 and 30 weeks, respectively. In Fig. 2(a), only pitting



(a)



(b)

FIG. 3. Ti 2p and O 1s XPS spectra acquired from the surface of the NiTi specimens after immersion in the 0.9% NaCl physiological solution for 2, 5, 8, 12, and 30 weeks.

pores can be found on the surface while the dark area around the pitting pores disappear after 12 weeks of immersion. Figure 2(b) shows a different surface morphology after immersion for 30 weeks. The entire surface is covered by a uniform and dense film instead of many pitting holes, although some tiny pores still exist on the surface. Undoubtedly, this film resulting from prolonged corrosion by aggressive Cl^- is mainly composed of titanium oxides. Our results show proof that the corrosion-induced titanium oxide layer grows and cures the pitting pores on the NiTi surface if the immersion time is long enough.

To obtain more detailed information on the constituents of the corrosion-induced titanium oxide layer on the NiTi specimens, XPS is used to examine the elemental chemical states. Figures 3(a) and 3(b) display the surface Ti 2p and O 1s spectra of the NiTi specimens after different immersion time. As shown in Fig. 3(a), the Ti 2p spectrum exhibits two main peaks with binding energies of about 465.0 eV ($\text{Ti}^{4+} 2p_{1/2}$) and 459.2 eV ($\text{Ti}^{4+} 2p_{3/2}$), indicating that titanium on the surface is mainly in the form of titanium oxide.²³ The intensity of both the

465.0 eV ($\text{Ti}^{4+} 2p_{1/2}$) and 459.2 eV ($\text{Ti}^{4+} 2p_{3/2}$) peaks increases with immersion time in the 0.9% NaCl solution. The O 1s spectra are shown in Fig. 3(b). The main peak has a binding energy of 530.8 eV, which can be assigned to oxygen present in metal oxides,²² and the peak intensity increases with immersion time. This indicates that titanium oxide becomes the dominant corrosion product on the surface of the NiTi specimens as the immersion time increases. The peak broadening observed from the O 1s peak after immersion for 12 weeks may be attributed to absorption of $-\text{OH}$ and $-\text{CO}$ groups on the surface.⁸

Figure 4 shows the XPS depth profiles of the corroded NiTi specimens compared with the mechanically polished one. The thickness of the corrosion-induced titanium oxide layer can be estimated from the depth profiles at which the oxygen signal drops to one-half of its surface value.²⁴ Taking a sputtering rate of 7.8 nm/min, the thickness of the native oxide film on the mechanically polished samples is about 4 nm, which is within the range reported for NiTi in the literature.^{1,14,25} On the other hand, the thicknesses of the titanium oxide films on the NiTi specimens after immersion for 12 and 30 weeks are about 31

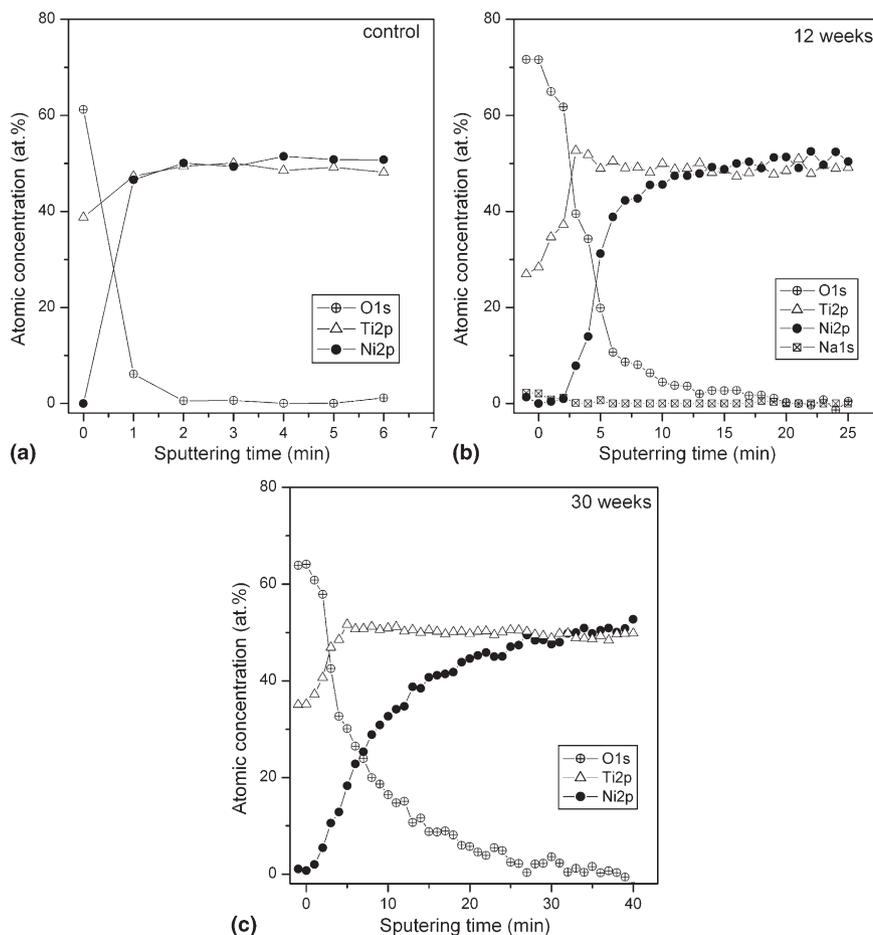


FIG. 4. XPS depth profiles obtained from the surface of the NiTi SMAs. (a) Mechanically polished NiTi without immersion and NiTi specimens immersed in the 0.9% NaCl physiological solution for (b) 12 weeks and (c) 30 weeks.

and 40 nm, respectively. It should be noted that the thickness of the corrosion product layer does not increase linearly with time. The growth rate decreases as corrosion proceeds, indicating that the as-formed titanium oxide layer may impede further corrosion of the NiTi SMAs in the NaCl solution.

High-resolution XPS measurements of both the Ti and Ni binding energy as a function of the sputtering time were conducted on the NiTi SMAs after immersion for 30 weeks. The Ti 2p spectra are corrected for background and scattering using the Shirley method.²⁶ Deconvolution of the Ti 2p spectra is accomplished using four oxidation states, Ti^{4+} , Ti^{3+} , Ti^{2+} , and Ti^0 , with corresponding $2p_{3/2}$ peaks at 459.2, 457.5, 455.3, and 454.5 eV, and $2p_{1/2}$ peaks at 465.0, 463.5, 461.5, and 460.6 eV.²³ It can be observed from Fig. 5(a) that Ti at the outmost surface (curve 1) is essentially composed of Ti^{4+} , Ti^{3+} , and Ti^{2+} corresponding to TiO_2 , Ti_2O_3 , and TiO . There is no Ti^0 as the intermetallic NiTi state, indicating that the outermost surface layer is mainly composed of titanium oxides. Deeper into the NiTi substrate, two peaks Ti^0 $2p_{3/2}$ and $2p_{1/2}$ corresponding to intermetallic NiTi state emerge and change gradually into the dominant peaks. After sputtering for 30 min, the main chemical state of Ti is intermetallic NiTi. It is noted that there is no NiO observed on the corroded surface according to our XPS results in spite of the fact that NiTi contains a large amount of Ni. This is caused by the difference in the oxygen affinity between Ti and Ni. The heats of formation of TiO_2 and NiO are approxi-

mately -956 and -241 kJ/mol, respectively,²⁷ implying that means Ti is easily oxidized whereas Ni is more difficult to oxidize. The titanium oxide layer is composed of transient Ti oxides, TiO , and Ti_2O_3 , because of insufficient oxidation at low temperature. The Ti–oxygen bond is more stable, and the resulting oxides do not easily dissolve into the solution after corrosion, which has been observed from H_2O_2 oxidized NiTi alloy.²⁸

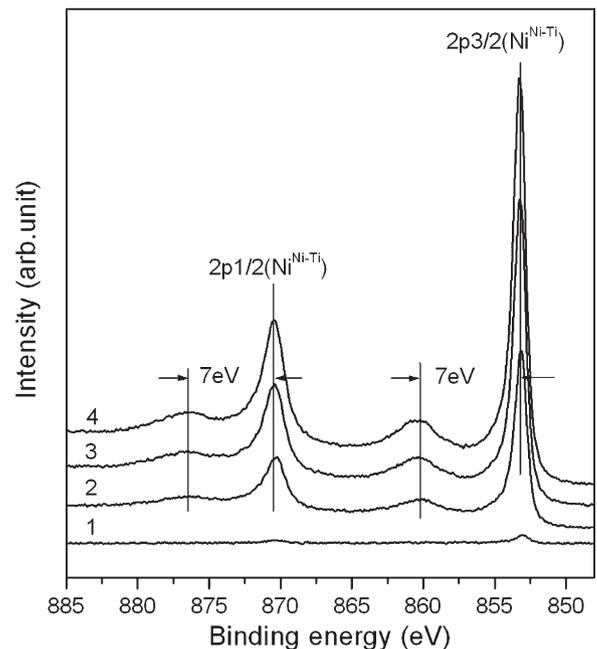


FIG. 6. Ni 2p XPS spectra obtained from the surface of the NiTi specimens immersed in the 0.9% NaCl physiological solution for 30 weeks at different sputtering times: (1) 0 min, (2) 5 min, (3) 10 min, and (4) 30 min.

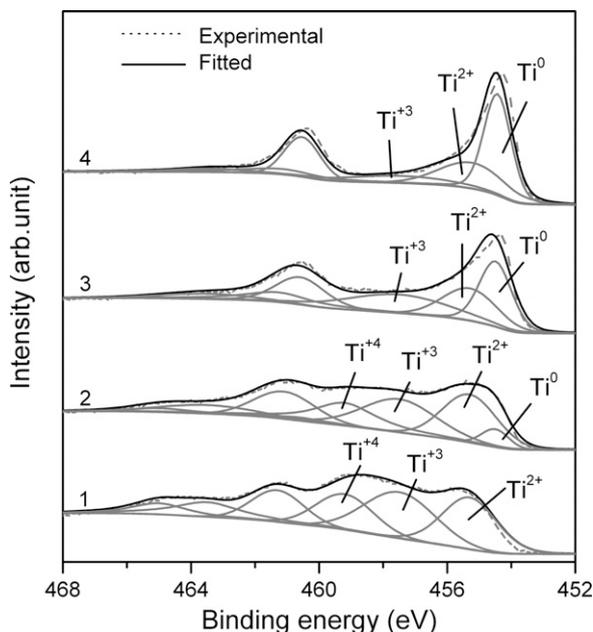


FIG. 5. High-resolution Ti 2p XPS spectra acquired from the surface of NiTi specimens immersed in the 0.9% NaCl physiological solution for 30 weeks at different sputtering times: (1) 0 min, (2) 5 min, (3) 10 min, and (4) 30 min.

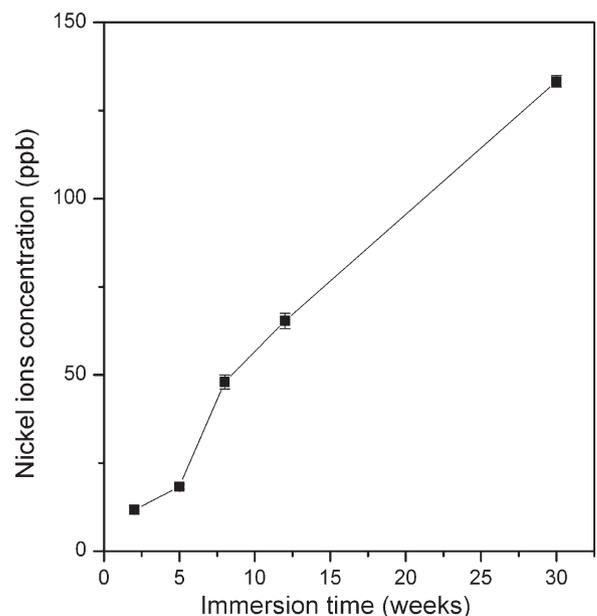


FIG. 7. Concentration of nickel ions released from NiTi specimens to the NaCl solution as a function of immersion time.

Figure 6 shows the high-resolution XPS collections of Ni 2p spectra. Nearly no Ni peaks can be observed from the outermost surface layer (curve 1), implying that Ni ions are lost to the solution during the corrosion process giving rise to a nickel-free surface on the corroded NiTi specimen. At a greater depth, the Ni 2p spectra consist of two major peaks near 853.8 and 870.5 eV, which correspond to Ni⁰ 2p_{3/2} and 2p_{1/2} in the intermetallic NiTi state. Two small peaks separated from the main peaks by ~7 eV are attributed to the satellite structure of the main peaks.²⁹ The changes in the Ni 2p spectra are consistent with those observed from the Ti 2p spectra.

TABLE II. Concentration (ppb) of nickel leached from NiTi specimens immersed in the 0.9% NaCl physiological solution for 2, 5, 8, 12, and 30 weeks, respectively.

	Immersion time (weeks)				
	2	5	8	12	30
Nickel ion concentration (ppb)	11.7 ± 1.3	18.3 ± 0.7	48.0 ± 2.0	65.3 ± 2.2	133.3 ± 1.7

To assess the release of nickel ions from the NiTi specimens into the NaCl solution, ICPAES was adopted. Figure 7 shows the amounts of Ni released from NiTi specimens to the NaCl solution as a function of immersion time, and Table II gives the Ni concentrations. The amounts of Ni ions increase with immersion time. It can be observed that the slope increases after immersing for 5 weeks and then decreases after 8 weeks. Since the slope of the curve in Fig. 7 is equal to the Ni release rate, it is clear that the nickel release rate varies with immersion time and is higher between 5 and 8 weeks. It can thus be inferred that the change in the release rate is related to the corrosion defects such as pitting pores on the NiTi specimens. In the early immersion period, nickel ions are released gradually into the surrounding solution. As the immersion continues, corrosion results in the formation of pitting pores that promote nickel release, leading to a high nickel release rate. However, if the immersion time is long enough, the corrosion-induced titanium oxide layer seals the pitting pores on the NiTi surface, thereby reducing the Ni release rate.

According to the aforementioned experimental results, a mechanism is proposed and the schematic diagram is

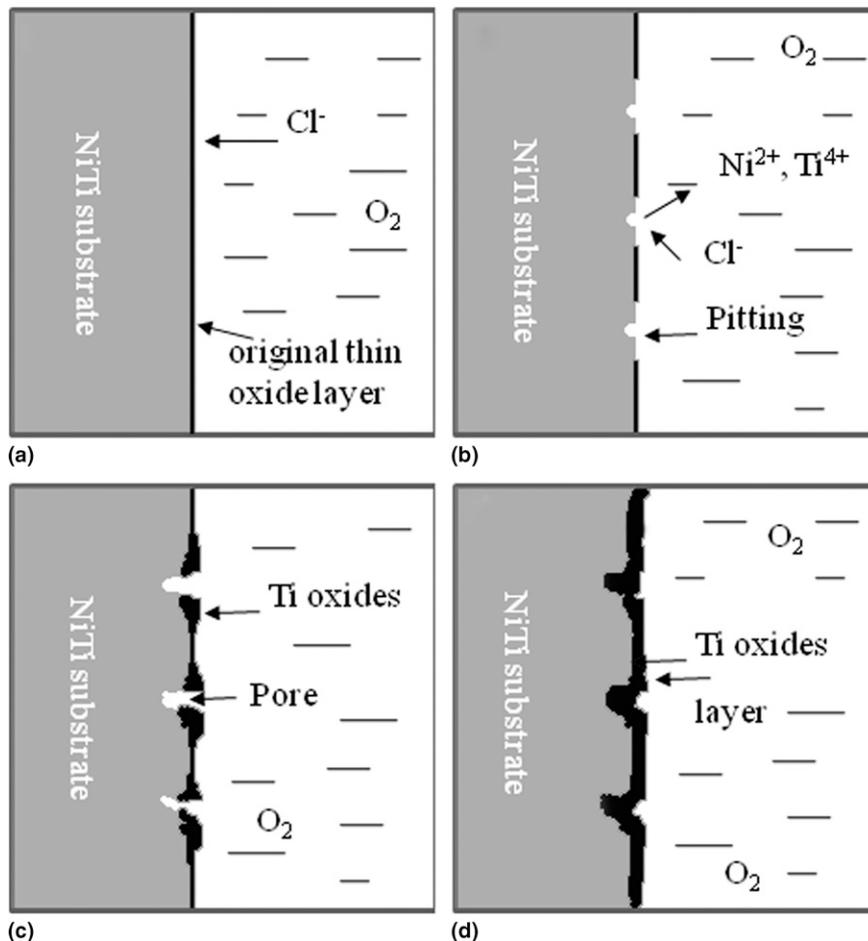


FIG. 8. Corrosion product layer formation on NiTi SMAs in the NaCl solution.

illustrated in Fig. 8. Upon exposure to the NaCl solution, the original thin oxide layer on the NiTi specimens are vulnerable to attack by Cl^- as shown in Fig. 8(a). As shown in the SEM image in Fig. 1(a), early pitting corrosion takes place on many sites on the surface. This process is illustrated in Fig. 8(b). Early corrosion process takes place on surface defects initially and involves the breakdown of original thin oxide layer and Ni–Ti bond after attack by Cl^- . This process releases nickel ions. As time elapses, the pitting sites propagate and noticeable pores form on the surface as confirmed by the SEM result in Fig. 1. Nickel released from the bulk materials into the solution reduces the local concentration of nickel in the materials. Meanwhile, the remaining titanium reacts with dissolved oxygen in the solution to form titanium oxides around the pitting sites and pores, as shown in Fig. 8(c). As the corrosion process proceeds, the titanium oxide layer grows and propagates over the entire surface of the NiTi specimen, as shown in Fig. 2(b). Moreover, the pores formed during corrosion are possibly blocked by the titanium oxides formed inside the pores. As a result, a uniform and dense oxide layer is formed on the NiTi SMAs, as shown in Fig. 8(d).

After the titanium oxide layer has formed on the entire surface of the NiTi specimen, it will serve as a passivation film to retard the corrosion process. Many studies have demonstrated that the titanium oxide film on the NiTi alloy serves as a barrier against outward diffusion of metal ions from the bulk alloy. Although titanium oxides possess good biocompatibility, the biological properties also depend on the quality of the titanium oxides film, e.g., surface defects, film bonding strength to the NiTi substrate, and so on. In this study, the corrosion product layer is found to consist of mainly titanium oxides formed during the corrosion process in the 0.9% NaCl physiological solution. These findings suggest that NiTi shape memory alloys can improve their own biocompatibility by the corrosion-induced titanium oxides formed on its surface during corrosion.

IV. CONCLUSION

Mechanically polished NiTi alloy samples are subjected to corrosion in a 0.9% NaCl physiological solution for 2, 5, 8, 12, and 30 weeks. The corrosion product layers on the NiTi specimens are characterized by SEM, EDS, and XPS. The following conclusions can be drawn:

(1) Pitting corrosion initiates on the surface of NiTi specimens in the NaCl solution. As the immersion time increases, pitting corrosion propagates onto the entire surface of the materials.

(2) During the corrosion process, nickel ions are released from the bulk materials into the solution. Meanwhile, the remaining titanium reacts with oxygen to form titanium oxides in the corrosion product layer.

(3) The corrosion product layer is composed of TiO_2 as well as other sub Ti oxides, TiO , and Ti_2O_3 and is relatively depleted of Ni.

(4) The growth rate of the titanium oxide film decreases with corrosion time, and the film impedes further corrosion.

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